in addition to other products, are formed) studied in recent years by Japanese workers. However, the $RHgX-CO_2(CO)_8$ reaction differs considerably from these reactions which involve more reactive organometallic reagents in over-all mechanism and in that it proceeds cleanly and does not give a mixture of products when carried out properly.

Our studies of this reaction and its extensions, with the aim of defining more fully its scope and its mechanism, are continuing.

Acknowledgments. The authors are grateful to the U. S. Army Research Office (Durham) for generous support of this work. This investigation was supported in part by Public Health Service Fellowship 5-F1-GM-32,971 (to R. J. S.).

(13) National Institutes of Health Predoctoral Fellow.

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Received November 22, 1967

Additions and Corrections

Halomethyl-Metal Compounds. XII. The Action of Sodium Iodide on Phenyl(trihalomethyl)mercury Compounds. A New Method of Dihalocarbene Generation. [*J. Am. Chem. Soc.*, **89**, 959 (1967)]. By DIETMAR SEYFERTH, MICHAEL E. GORDON, JEFFREY YICK-PUI MUI, and JAMES M. BURLITCH, Department of Chemistry, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139.

In Table I, the last entry in the third column [Product (yield, %)] should read: 2,2-Dichlorocyclopropyl acetate (14) and $CH_3CH(CCl_3)O_2CCH_3$ (7).

A Reinterpretation of the Nuclear Magnetic Resonance Spectrum of Cyclobutene [*J. Am. Chem. Soc.*, **89**, 2047 (1967)]. By E. ALEXANDER HILL, Department of Chemistry, University of Wisconsin—Milwaukee, Milwaukee, Wisconsin 53201, and JOHN D. ROBERTS, Gates and Crellin Laboratories of Chemistry, California Institute of Technology, Pasadena, California 91109.

On page 2049, the legend to Figure 1, b and c, should read: b, vinyl hydrogens of a cyclobutene mixture of -2,2,3-d_3 and -1,3,3-d_3; c, methylene hydrogens of the same mixture as in b.

Inversion Barriers of Pyramidal (XY_3) and Related Planar (=XY) Species. [*J. Am. Chem. Soc.*, **89**, 3396 (1967)]. By G. W. KOEPPL, D. S. SAGATYS, G. S. KRISHNAMURTHY, and SIDNEY I. MILLER. Department of Chemistry, Illinois Institute of Technology, Chicago, Illinois 60616.

On page 3400, column 2, line 14 add the following sentence: Configurationally stable rather than optically active phosphites are known (ref 20f). On page 3401, column 2, line 8 [6.1.0] should read [4.1.0]. On page 3402, in the Isotopes Effects section: deuterium ions should read deuteronium ions. On page 3404, column 1 delete the last paragraph. In Table I, for $SnCl_3^-$, $\nu_1 = 297$; for PT_3, $k_\delta/l^2 = 3.0825$; for IO_3^-, $\nu_2 = 390$. In Tables I, II, and III, delete all entries for NF_3, NCl_3, and PI_3. In Table III, delete the last entry for =NF.

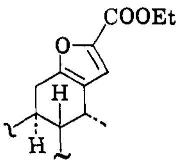
We now have a program in which any set (or sets) of force constants can be introduced into the calculation to replace those generated from the input data. By this technique, we generated V_i (kcal/mole) as follows: NF_3, 56–59; PF_3, 29–66; PI_3, 43; AsF_3, 36–41; AsI_3, 25–40. Five or six trials with arbitrary force constants for PF_3 gave equally arbitrary values of V_i = 29–66 kcal/mole.

The Synthesis of 3-(2′-Deoxy-D-ribofuranosyl)adenine. Application of a New Protecting Group, Pivaloyloxymethyl (Pom) [*J. Am. Chem. Soc.*, **89**, 5439 (1967)]. By MALCOLM RASMUSSEN and NELSON J. LEONARD. Department of Chemistry and Chemical Engineering, University of Illinois, Urbana, Illinois 61801.

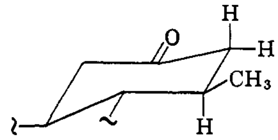
On pages 5439, 5441, 5442, and 5444, the preferred name for **2** is 3,5-di-O-(*p*-chlorobenzoyl)-2-deoxy-D-ribofuranosyl chloride.

Total Synthesis of Racemic Methyl Vinhaticoate. [*J. Am. Chem. Soc.*, **89**, 5497 (1967)]. By T. A. SPENCER, R. M. VILLARICA, D. L. STORM, T. D. WEAVER, R. J. FRIARY, J. POSLER, and P. R. SHAFER. Department of Chemistry, Dartmouth College, Hanover, New Hampshire 03755.

On page 5498 formulas **17** and **13** should appear as shown here, rather than as published.



17



13

One Electron Transfer Oxidation of 7,12-Dimethylbenz[*a*]anthracene, a Model for the Metabolic Activation of Carcinogenic Hydrocarbons [*J. Am. Chem. Soc.*, **89**, 5508 (1967)]. By JOSEF FRIED and DOROTHY E. SCHUMM. Ben May Laboratory for Cancer Research and Department of Chemistry and Biochemistry, University of Chicago, Chicago, Illinois 60637.